Launching Rockets using Secret Fuel

You will make a rocket using plastic bottles with volumes between 1 L and 4 L. You will fill the bottles with a mixture of hydrogen gas and oxygen gas in an appropriate ratio. When ignited, the reaction will produce water gas at a very high temperature. The hot molecules will generate a propulsive force (or the bottle will simply blow up).

1. The first thing to determine is the appropriate ratio of hydrogen and oxygen to use in your rocket. You need an explosive mixture to launch your rocket. A mixture too rich in \( \text{H}_2 \) will burn quietly like a Bunsen burner instead of igniting explosively. A mixture too rich in \( \text{O}_2 \) will explode, but weakly. A mixture that is just right will produce maximum power for your rocket. The unbalanced equation for the propulsion reaction is as follows:

\[
\text{H}_2(g) + \text{O}_2(g) \rightarrow \text{H}_2\text{O}(g)
\]

If you balance the equation shown above, you will know the correct mole ratio of \( \text{H}_2 \) to \( \text{O}_2 \) needed to achieve the maximum propulsive force.

2. Use a 1-L graduated cylinder to accurately measure the total volume of your bottle. Using the mole ratio of \( \text{H}_2 \) to \( \text{O}_2 \) that you discovered from your balanced equation, you can figure out how much \( \text{H}_2 \) and how much \( \text{O}_2 \) you will need to fill the plastic bottle. In the space below, sketch a picture of your bottle (label total volume) and draw a line showing what fraction of the bottle you will fill with \( \text{H}_2 \) and what fraction you will fill with \( \text{O}_2 \). Convert these fractions into actual volumes of \( \text{H}_2 \) and \( \text{O}_2 \) (in liters). Discuss how the mole ratio compares to the volume ratio that you sketch in your picture.

3. Use the ideal gas law and your picture above to calculate how many moles of \( \text{H}_2 \) you will need, and how many moles of \( \text{O}_2 \) you will need in your bottle. Also give a one-sentence explanation for why 22.4 L/mole is not a valid measurement to use for this experiment.
You just determined how much \( \text{H}_2 \) and \( \text{O}_2 \) you need for your rocket, but you will have to generate these gases by running chemical reactions.

**This part discusses production of \( \text{O}_2 \) gas.**

4. \( \text{O}_2 \) can be generated from the decomposition of \( \text{H}_2\text{O}_2 \). The unbalanced equation is:

\[
\text{H}_2\text{O}_2(\ell) \rightarrow \text{H}_2\text{O}(\ell) + \text{O}_2(\text{g}) \quad \Delta H = -190 \text{ kJ/mol}
\]

The reactant in this reaction, hydrogen peroxide, is thermodynamically unstable and can rearrange its atoms to form water and oxygen gas. This reaction is extremely slow at room temperature unless a catalyst is added. You will use potassium iodide (KI) as a catalyst to speed up this reaction. When running the actual reaction, a small scoop of the catalyst (about 0.7 grams) will be “packaged” in weighing paper and then added to the \( \text{H}_2\text{O}_2 \). You will place the stopper on the reaction flask and then shake it to allow KI to come out of its package. This method should enable you to collect \(~100\%\) of the gas created in this reaction.

a) Balance the equation shown above and use stoichiometry to calculate the mass of \( \text{H}_2\text{O}_2 \) needed to generate enough \( \text{O}_2 \) for your rocket. (Remember, you calculated the \( \text{O}_2 \) quantity in step #3.)

b) The \( \text{H}_2\text{O}_2 \) that you will use to generate \( \text{O}_2 \) is a liquid, which is a solution of \( \text{H}_2\text{O}_2 \) and \( \text{H}_2\text{O} \). The solution contains 10\% \( \text{H}_2\text{O}_2 \) by mass, the other 90\% is \( \text{H}_2\text{O} \). Use your value from (a) to calculate how many grams of the solution you will need. You can draw a picture of what the 10\% solution of \( \text{H}_2\text{O}_2 \) will look like to help with this calculation.

c) Assume the 10\% \( \text{H}_2\text{O}_2 \) solution has a density of 1.0 g/mL. Convert the grams you calculated in (b) into milliliters of solution.
d) Refer to the sketch that represents the energy diagram of the uncatalyzed decomposition of H₂O₂.
   i. In the sketch, label the following: H₂O₂, H₂O + O₂, activation energy (Ea), and ΔH.
   ii. Below, sketch an energy diagram for the catalyzed decomposition of H₂O₂. Then, to the right, explain how the catalyst helps the reaction proceed at room temperature.

   ![Uncatalyzed Reaction](image)

   ![Catalyzed Reaction](image)

**This part discusses the production of H₂.**

5. Hydrogen can be generated in many ways. You will use the method shown below:
   \[ \text{Ca(s) + H}_2\text{O(l)} \rightarrow \text{H}_2(g) + \text{Ca(OH)}_2(s) \quad \Delta H = \_\_\_\_\_\_\_\_\_ \]

a) Balance the equation shown above.

b) Calculate ΔH for the reaction (in kJ/mole) with the following information: the reaction is exothermic and produces 10.4 kJ for each gram of calcium that reacts. Write the calculated value in the space provided above.

c) Use stoichiometric calculations to determine the mass of Ca and volume of H₂O needed to produce enough H₂ for your rocket. Remember, do you know how much H₂ you need?
d) The amount of H₂O you have calculated is the amount that will actually react with the Ca that you will weigh out. This is not a smart way to run this reaction. Circle two of the choices below that express reasons why it is a good idea to add excess H₂O when you carry out this reaction. When you run the reaction, you will want to use a large excess (as large as your reaction flask permits). To the right, sketch what your flask should look like if you use a large excess of H₂O.
   i. extra water will create extra hydrogen
   ii. extra water will create more heat in the reaction flask
   iii. extra water will keep the reaction flask cool
   iv. extra water will reduce the amount of air that goes into the rocket

e) Calcium is a very reactive metal. Even while inside its jar, the outer surface of calcium is oxidized to some extent (forming CaO). This reduces the amount of actual calcium in the lumps that you will use. How can you correct for this problem?

On page five, see the instructions for how to prepare and launch your rocket.

After you launch your rocket, answer these questions:
6. Discuss the role of the ignition spark in the combustion reaction that launched the rocket. Sketch an energy diagram in the space to the right and use this diagram to discuss the term “activation energy.”

7. Why does the bottle actually lift off? (Note: the ignition of the H₂/O₂ mixture generates very hot gas molecules with lots of kinetic energy. What do you think these hot molecules do?)
Step-by-step instructions for preparing and launching your rocket

1. You will need the following equipment to fill your bottle with hydrogen and oxygen gas:
   a) two reaction flasks—small and large
   b) one-holed stoppers that fit each flask
   c) rubber hose with glass dropper tip
   d) water trough (or bucket)
   e) the bottle that is to be your rocket (with lid)

2. Put tap water in the trough until it is approximately half-full.

3. Fill your rocket bottle completely with water, cap it, and invert it in the trough. You can leave the cap on the bottle until you are ready to load the first gas (O₂).

4. Measure the appropriate volume of 10% H₂O₂ and place it in your small reaction flask. Use of the small flask (as opposed to the large flask) will minimize the amount of air that goes into your rocket. The appropriate volume of H₂O₂ for your rocket should be found in section 4(b).

5. Use weighing paper to measure out approximately 0.7 grams of KI, the catalyst.

6. Fold your weighing paper around the catalyst so that it makes a loose package, like an envelope. This packaging of the catalyst will delay the decomposition reaction until you have a stopper on the reaction flask. The goal is to get 100% of your O₂ transferred to the rocket, without losing the initial “burst” of gas that will be produced when catalyst hits H₂O₂.

7. Take the lid off your bottle that is sitting in the trough. Feed the end of the rubber tube into the mouth of the bottle. Attach the other end of the rubber tube (with glass dropper tip) to the stopper that fits in the flask. You are now ready to transfer gases.

8. Drop the package of catalyst into the reaction flask, which already contains the measured volume of H₂O₂. Place the stopper on the flask and shake the flask so that the catalyst comes out of its envelope. O₂ should begin bubbling out of the flask and into the rocket bottle.

9. Continue the collection of O₂ until the reaction in the flask stops. When no more bubbles are being transferred, remove the rubber hose from the rocket bottle. You may wish to replace the cap on the rocket bottle at this time. Note: Your bottle should be 1/3-filled with O₂!

10. Pour the remaining liquid in the reaction flask down the drain. Take the paper out of the reaction flask and put it into the trash can, not the sink.

11. Fill your large reaction flask to the 150-mL mark with tap water. This water will be used in the reaction that generates H₂.

12. Find the quantity of calcium metal that is appropriate for your rocket (section 5(b)). It is recommended that you use approximately 0.3 grams more calcium than your calculated quantity. This extra calcium will ensure that you fill your bottle completely with gas. Weigh out the recommended quantity of calcium metal on weighing paper.

13. Make sure your hose and stopper are set up so that you can transfer gas from the reaction flask to the rocket bottle.
14. **Read this entire paragraph before doing any of the steps involved.**

Drop the calcium metal into the reaction flask (that already should contain 150 mL of tap water). You do not need to “package” the calcium metal like you did the KI. Quickly put the stopper on and allow the calcium to react until the rocket bottle is completely filled with gas. As soon as the bubbles start coming out the bottom of the rocket bottle, you should remove the hose from the mouth of the bottle. Your goal is to fill the bottle with gas—adding more H₂ after this point will make the fuel mixture too rich in H₂! Note: As the rocket bottle gets filled with H₂, it will become very light, which will cause it to float in water. One member of your group should firmly hold the bottle so it does not tip over.

15. If your rocket bottle is not completely filled with gas (i.e. if there is some water still inside the bottle), you may either generate a bit more H₂ by throwing a couple more rocks of calcium into the reaction flask to displace the remaining water or you can physically lift up the bottle and let the water drain out of the bottle (recommended only if you have a very small amount of water left in the bottle).

16. Once your rocket bottle is completely filled with gas, you should lift it out of the water trough and cap the bottle tightly. Note: the gas mixture in the rocket bottle is lighter than air; therefore you must hold the bottle in an inverted orientation to avoid losing all your fuel!!!!

17. Clean up the reaction flask by pouring the white liquid down the drain and rinsing the flask with water. If the flask is still hazy with white residue, you can either scrub it or use some wash acid to clean out the flask.

18. Keep your safety glasses on even though you think you are “done” with the experimental part of the lab. Your rocket is an explosive device and you must continue to think about safety precautions.

19. Your rocket is now ready to launch. You are responsible for bringing your rocket bottle and your igniters to the designated launch area of the football field. You are also responsible for cleaning up the launch area after the launches are finished.